

## Stoichiometry Chapter Test

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### Chapter 3: Chemical Reactions and Reaction Stoichiometry ...

Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers:  $2 \text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow 2 \text{NO}_2 + 3 \text{H}_2\text{O}$ . The question asks about the amount of oxygen reacting with ONE mole of ammonia, thus cut the  $\frac{7}{2}$  (3.5) of oxygen in half to 1.75 2. c Balance: CH

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### Chapter 9 Reading Guide Stoichiometry

Stoichiometry Practice Test 4. How many liters of  $\text{O}_2$  at STP will react with 8.7 grams of  $\text{C}_2\text{H}_4$  to form  $\text{CO}_2$  and  $\text{H}_2\text{O}$ ?  $\text{C}_2\text{H}_4 + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$  A. 20.88L B. 2.32 L C. 22.4 L D. 6.54 L 5. Using the equation in question number 4, what is the mole ratio of carbon dioxide to oxygen? A. 1:2 B. 3:2 C. 2:2 D. 2:3 6.

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Chapter 5. Stoichiometry and the Mole. Stoichiometry Learning Objectives. Define stoichiometry. ... Test Yourself. How many molecules of H<sub>2</sub> are needed to react with 29 molecules of N<sub>2</sub> to make ammonia if the balanced chemical equation is N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub>? Answer. 87 molecules.

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