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a. 2.89 g of NaCl dissolved in 0.159 L of water (density of water is 1.00 g/mL) 0.311 molal NaCl b. 1.80 mol KCl in 16.0 mol of H₂O 6.25 molal KCl c. 13.0 g benzene, C₆H₆ in 17.0 g CCl₄ 9.80 molal C₆H₆

Practice Problems: Solutions (Answer Key)

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Determine the mass of calcium hydroxide, Ca(OH)₂, produced when calcium carbide reacts with 0.64 g of water, H₂O, according to the following equation: CaC₂ + 2H₂O (Ca(OH)₂ + C₂H₂. 0.64 g x 1 moles H₂O = 0.0355 mole H₂O. 18 g. 0.0355 H₂O mole x 1 mole Ca(OH)₂ = 0.01778 mole Ca(OH)₂. 2 mole H₂O

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Balancing Equations: Practice Problems

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2-2 Practice Problems, page 11 . Answers marked with an asterisk denote . additional practice problems that appear in the . Teacher's Edition. *1. 320 K· 2. 1273 K . 3. 253 K 4. 173 K *5. -61°C 6. -241°C 7. 3850°C 8. 28.0 °C 9. 293 K 10. 1337 K . 2-2 Apply, page 12 . 1. H.e added 173 to the Celsius temperature ...

2 Answer - Weebly

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